

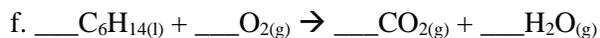
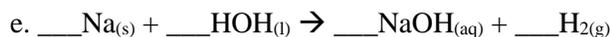
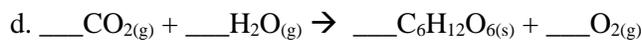
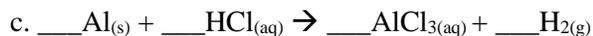
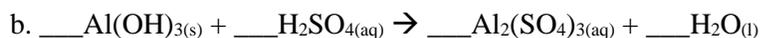
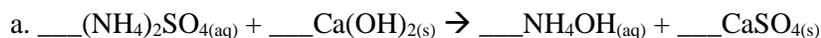
Questions to be completed and submitted to work towards “credit earned” in Chemistry 112.
Please email a picture or scanned copy of your work to Mrs. Arsenault.

Material covered April 14 – 17

1. Convert

- 1.20×10^{24} atoms of sodium to moles
- 13.2 moles of carbon dioxide to molecules
- 10.0 moles of oxygen gas to molecules
- 6.50 moles of magnesium to atoms
- 6.63×10^{23} molecules of magnesium acetate to moles
- 3.33×10^{25} molecules of iodine to moles

2. Balance the following chemical equations



Material covered April 20 - 24

1. Calculate:

- Moles of CH_3COOH in 136.4 grams
- Moles of $\text{Ca}(\text{ClO}_3)_2$ in 25.7 grams
- Grams of $\text{Al}(\text{NO}_3)_3$ in 0.45 moles
- Grams of SiO_2 in 1.75 moles