

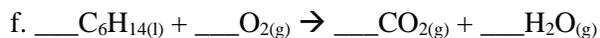
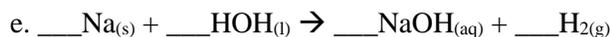
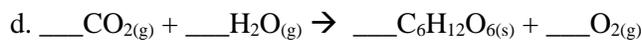
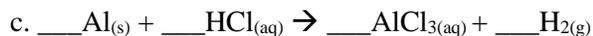
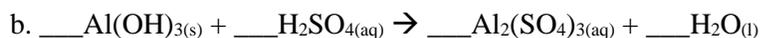
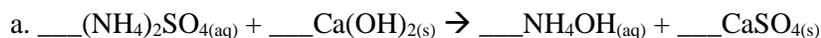
Questions to be completed and submitted to work towards "credit earned" in Chemistry 112.  
Please email a picture or scanned copy of your work to Mrs. Arsenault.

Material covered April 14 – 17

1. Convert

- $1.20 \times 10^{24}$  atoms of sodium to moles
- 13.2 moles of carbon dioxide to molecules
- 10.0 moles of oxygen gas to molecules
- 6.50 moles of magnesium to atoms
- $6.63 \times 10^{23}$  molecules of magnesium acetate to moles
- $3.33 \times 10^{25}$  molecules of iodine to moles

2. Balance the following chemical equations



Material covered April 20 - 24

1. Calculate:

- Moles of  $\text{CH}_3\text{COOH}$  in 136.4 grams
- Moles of  $\text{Ca}(\text{ClO}_3)_2$  in 25.7 grams
- Grams of  $\text{Al}(\text{NO}_3)_3$  in 0.45 moles
- Grams of  $\text{SiO}_2$  in 1.75 moles