

Questions to be completed and submitted to work towards credit earned in Chemistry 112.
Please email a picture or scanned copy of your work to Mrs. Arsenault.
All credit earned material must be passed in by June 12.

Material covered June 1 – 5

1. Write balanced net ionic equations for the following:
 - a. $\text{Pb}(\text{NO}_3)_2(\text{aq}) + \text{Na}_3\text{PO}_4(\text{aq}) \rightarrow \text{NaNO}_3(\text{aq}) + \text{Pb}_3(\text{PO}_4)_2(\text{s})$
 - b. $\text{NaOH}(\text{aq}) + \text{CaCl}_2(\text{aq}) \rightarrow \text{NaCl}(\text{aq}) + \text{Ca}(\text{OH})_2(\text{s})$
 - c. $\text{Pb}(\text{NO}_3)_2(\text{aq}) + \text{NaCl}(\text{aq}) \rightarrow \text{PbCl}_2(\text{s}) + \text{NaNO}_3(\text{aq})$
2. Identify the precipitate formed when the following ionic compounds react:
 - a. $\text{AgNO}_3(\text{aq}) + \text{H}_2\text{S}(\text{aq}) \rightarrow$
 - b. $\text{Pb}(\text{NO}_3)_2(\text{aq}) + \text{NaOH}(\text{aq}) \rightarrow$
 - c. $\text{KI}(\text{aq}) + \text{AgNO}_3(\text{aq}) \rightarrow$
3. Will a precipitate form when the following compounds react? If so, write a net ionic equation for the formation of the equation.
 - a. $\text{AgNO}_3(\text{aq}) + \text{Na}_2\text{SO}_4(\text{aq})$
 - b. $\text{NH}_4\text{Cl}(\text{aq}) + \text{Ba}(\text{NO}_3)_2(\text{aq})$
 - c. $\text{CaCl}_2(\text{aq}) + \text{K}_2\text{SO}_4(\text{aq})$