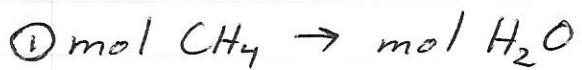
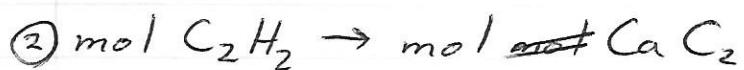


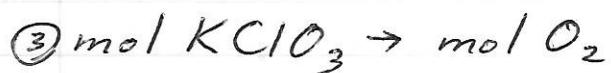
Mole ratios
Practice problems



$$5.86 \text{ mol CH}_4 \times \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol CH}_4} = 11.72 \text{ mol H}_2\text{O}$$



$$5.06 \text{ mol C}_2\text{H}_2 \times \frac{1 \text{ mol CaC}_2}{1 \text{ mol C}_2\text{H}_2} = 5.06 \text{ mol CaC}_2$$



$$8.55 \text{ mol KClO}_3 \times \frac{3 \text{ mol O}_2}{2 \text{ mol KClO}_3} = 12.83 \text{ mol O}_2$$



$$8.64 \text{ mol H}_2\text{O} \times \frac{3 \text{ mol NO}_2}{1 \text{ mol H}_2\text{O}} = 25.92 \text{ mol NO}_2$$

