Properties of ionic compounds

1. Most ionic compounds are crystalline solids at room temperature. This is a result of the ions being arranged in repeating three-dimensional patterns.
2. Ionic compounds usually have high melting points due to the strong attractive forces within the compound.
3. Ionic compounds can conduct electric current when melted or dissolved in water. The crystalline structure breaks down allowing the anions to migrate to the anode and the cations to migrate to the cathode. The movement of the ions allow electric current to flow.