**Chapter 9**

**Chemical Names and Formulas Assignment**

**Answer the following questions in sentence forms**

**Section 1**

1. What are monatomic ions? Provide four examples (Value 3)
2. What is a cation and how are they formed? Provide four examples (Value 4)
3. What is an anion and how are they formed? Provide four examples (Value 4)
4. How do the ions formed by metals in groups 1A, 2A and 3A differ from the ions formed by nonmetals in group 5A, 6A and 7A? (Value 3)
5. How are charges of some transitional metals determined (Value 1)
6. Are ions formed by transition metals positive or negative? (Value 1)
7. When transition metals are dissolved in water their ions will form a color solution. What color will the following ions form when dissolved in water and when placed in a flame? Co3+, Cr3+, Fr3+, Ni2+ and Mn2+ (Value 5)
8. What does the prefix (-ous) and (-ic) represent? (Value 2)
9. What are the Classic names for the Cu+ and Cu2+ ion? What is the chemical formula for Cupric Chloride and Cuprous Fluoride? (Value 4)
10. Complete question 1 and 2 on p256. (Value 6)
11. What is a polyatomic ion? Can a polyatomic ion have a positive charge? If so provide an example. (Value 3)
12. List the four possible suffixes for ions. (Value 4)

**Choose two of the following topics and provide some basic information. (Value 10)**

1. Pharmacist is required to have a strong background in chemistry. Read page 256 for some basic information about a career in pharmacy. Access the web link at PHSchool.com and describe some basic information on a career on Pharmacy. Which universities in Atlantic Canada offer a BSc in Pharmacy?
2. Access the web link at PHSchool.com (p 256) for information on careers. Choose one of those careers and write a description.
3. Complete a compare and contrast chart comparing monatomic and polyatomic ions. Include information about charges and name endings.
4. Complete the “Writing Activity” section on p258 on sodium and potassium ions.

**Section 2**

Read p260 – 266

Complete the following questions: p265: 12-13 (Value 4) and p266: 14-19 (Value 15)

Complete the “Element and Handbook” section on p266

**Section 3**

Read p268 – p270

Complete the following questions: p270: p22-25

Complete the “Connecting and Concepts” section on p270